

8 (14)

What is the change in entropy of 570g of ice as it melts into water.

Solution

The change in entropy of the ice melting into water is

$$\Delta S = \frac{\Delta Q}{T}$$

where: ΔQ is the amount of heat.

T is the temperature in Kelvin

$\Delta Q = mL_f$, where L_f is the latent heat of fusion of ice = 333 J/g

$$T = 273 + 0 = 273 \text{ K.}$$

$$\Delta S = \frac{mL_f}{T} = \frac{333 \times 570}{273}$$

$$\Delta S = 695.2747 \text{ J/K}$$

9 (1b)

Forty moles of monatomic gas expand isobarically from volume 3.5 m^3 to volume 9.5 m^3 . If the gas has an initial temperature 45°C then,

(a) Find the final temperature.

Summary of question.

No. of moles $40 = n$

Initial volume $= 3.5 \text{ m}^3$

final volume $= 9.5 \text{ m}^3$

Initial temperature $T_i = 45^\circ\text{C}$

let final temperature be T_f

pressure $\Rightarrow P$.

$$\begin{aligned} \therefore PV_i &= nRT_i \\ \text{and } PV_f &= nRT_f \end{aligned} \left. \begin{array}{l} \text{pressure is same since} \\ \text{it is isobaric process.} \end{array} \right\}$$

(a) Find final temperature.

dividing both equation above

$$\frac{PV_i}{PV_f} = \frac{nRT_i}{nRT_f}$$

$$\therefore T_f = \frac{V_f T_i}{V_i} = \frac{9.5 \times (273 + 45)}{3.5} = \frac{9.5 \times 318}{3.5}$$

$$\begin{aligned} T_f &= \cancel{10573.5} \text{ K} \\ &= 863 - 273 = 590^\circ\text{C} \end{aligned}$$

9 (6)

b) The work done by expanding gas.

$$W = \int_{V_i}^{V_f} P dV \quad \text{since } P \text{ is constant.}$$

then

$$W = P \int_{V_i}^{V_f} dV.$$

$$W = P_{V_f} - P_{V_i}$$

$$W = nR\bar{T}_f - nR\bar{T}_i$$

$$W = nR(\bar{T}_f - \bar{T}_i)$$

$$W = 40(8.134) \left\{ \cancel{590} - 863 - 318 \right\}$$

$$= 40 \times 8.134 \times 545$$
$$= 177,321.2 \text{ J}$$

c) Heat gained or lost. Q

for constant pressure

$$Q = n C_p \Delta T \quad \text{where } C_p \text{ is molar heat capacity of constant pressure.}$$

$$C_p = \left(1 + \frac{f}{2}\right) R \quad f \Rightarrow \text{degree of free atom}$$

for monoatomic gas $f = 3$

$$\therefore C_p = \left(1 + \frac{3}{2}\right) R = \frac{5}{2} R$$

Continuation of # 9 (16) part C

$$Q = mc_p \Delta T = 40 \times \frac{5}{2} R (\bar{T}_f - \bar{T}_i)$$

$$= 40 \times \frac{5}{2} \times 8.314 \times (863 - 318)$$

$$= 831.4 \times 545$$

$$Q = 453113 \text{ Joule. (Heat gained)}$$

Question # 10 (16)

Copper pellets with total mass of 200g at a temperature 100°C dropped into a container holding 450g of water at 24°C . Find the equilibrium temperature of the system.

Solution

at equilibrium temperature

Heat lost = Heat gained

$$Q_{\text{lost}} = Q_{\text{gained}}$$

$Q = mc\Delta\theta$ where $m \Rightarrow$ mass, $c \Rightarrow$ specific heat capacity

$\Delta\theta \Rightarrow$ Change in temperature

specific heat capacity of water is $4200 \text{ J/kg/}^\circ\text{C}$

specific heat capacity of copper $385 \text{ J/kg/}^\circ\text{C}$

$$\frac{\text{Heat lost by copper}}{\text{water}} = \frac{\text{Heat gained by water}}{\text{water}}$$

$$Q_{\text{lost}} = 200 \times 385 \times (100 - \bar{T}_f) \quad (450 \times 4200 \times (\bar{T}_f - 24))$$

$$770000 - 77000\bar{T}_f = 1890000\bar{T}_f - 45360000$$

$$770000 + 45360000 = (1890000 + 77000)\bar{T}_f$$

$$\bar{T}_f = \frac{50130000}{1967000} = \underline{\underline{25.49^\circ\text{C}}} \Rightarrow \text{Final temperature.}$$

#12. (14)

A hot metal in the shape of cube whose edge has length 12cm is at temperature 600K. at what net rate does it radiate heat when it is placed in air at temperature 300K?

Solution

Applying Stefan Boltzmann's law, the heat emitted at a temperature T , having emissivity " e " and surface area A is given by

$$Q = \sigma e A T^4$$

where σ is Stefan Boltzmann's Constant

$$\sigma = 5.67 \times 10^{-8}$$

$$\text{Area} = 6 \times \left(\frac{12}{100}\right)^2 = 0.0864 \text{ m}^2.$$

$$e = 1$$

$$Q = 5.67 \times 10^{-8} \times 1 \times 0.0864 \times 300^4$$

$$= 39.681 \text{ J/sec}$$

11. (14)

A thin-walled bowl in the shape of hemisphere just barely floats in water. If it has a diameter of 10 cm, then what is its mass.

Solution

for equilibrium

buoyant force = weight

from Density = $\frac{\text{mass}}{\text{volume}}$

force = density of liquid \times volume $\times g$

then, weight = $mg = \text{mass of bowl} \times g$

$$\therefore \rho \times V \times g = mg$$

$$\therefore \rho \times V = m$$

density of water = 1000 kg/m^3

$$\text{Mass} = 1000 \times \frac{2}{3} \pi (5 \times 10^{-2})^3$$

$$r = 5 \text{ cm}$$
$$\pi = \frac{22}{7}$$

$$= 0.262 \text{ kg.}$$